Atoms, Elements and the Periodic Table Answers

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

Water 1. Which substance listed below is NOT an element?

atom 2. A(n) _____ is a basic unit on matter.

Neutrons 3. ____ are particles without electric charge in the atom's nucleus.

Bohr 4. Which scientist calculated what energy levels were around each element? We also use his model to draw the elements.

ion 5. A positively charge _ has more prrotons than electrons.

matter 6. Anything that takes up space and has mass is _____.

# of Neutrons 7. Isotopes of Carbon have different _____.

protons & neutrons 8. _____ are particles that are located in the atom's nucleus.

Periods 9. The horizontal rows in the Periodic Table are called the ___.

Groups & Families 10. The vertical columns of the Periodic Table are called the ___.

Protons 11. The atomic number refers to the number of _____.

Protons & Neutrons 12. The atomic mass is a combination of what two sub-atomic particles?

valence Electrons 13. The number of electrons on the outer energy level or shell is called _____.

7 14. Halogens all have ____ valence electrons.

Electrons 15. H_2O is a compound formed by sharing _____.

atomic # 16. The modern periodic table is arranged in order of increasing.

Group 17. Elements with similar characteristics are the same ____.

Energy levels 18. The horizontal rows on the periodic table tells us how many _____ the element has.

atoms 19. Elements are all alike in that they all contain __.

Like characteristics & = valence electrons 20. A family or group of elements have ____.

Electrons 21. An ion haas lost or gained ____.

Neutrons 22. An isotope has lost or gained ____.

Alkali Metals 23. Silvery solids with low densities and low melting points found in Group 1.

Alkaline Earth Metals 24. Another name for Group 2.

Transition Metals 25. Elements in group 3-12.

Noble Gases 26. Groups that has it's valence electrons filled.

Halogens 27. Groups that readily combine with Alkali Metals to form salts.

metal 28. A transition element

non- metal 29. Noble Gases

non- metal 30. Calcium

metalloids 31. Elements that show properties of both metals and non-metals.

metalloids 32. Boron & Silicon

2 electrons 33. 1st energy level

8 electrons 34. 2nd energy level

18 electrons 35. The 3rd energy level

32 electrons 36. 4th energy level

Bohr 37. Electrons travel definite paths

Rutherford 38. Gold Foil experiment

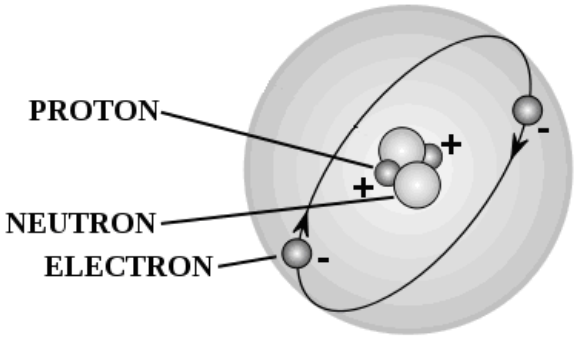
Thomson 39. Cathode Ray Tube

Democritus 40. Proposed Atoms

Dalton 41. Discovered the nucleus of the atom

Thomson 42. Plum pudding model

Dalton 43.Created the Atomic Theory



____ 44. Neutrons

____ 45. Protons

____ 46. Electrons

47. Fill in the missing information. Round all answers

Protons = 79
Electrons = 79
Neutrons = 118
Atomic Mass = 197
Atomic Number = 79

79
Au
Gold
196.967

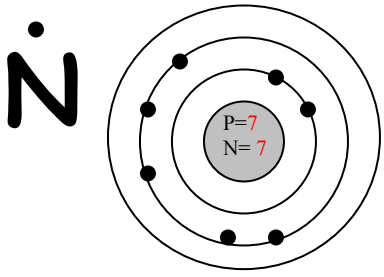
48. Fill in the missing information. Round all answers

Protons = 1
Electrons = 1
Neutrons = 0
Atomic Mass = 1
Atomic Number = 1

1
H
Hydrogen
1.008

49. Draw a Bohr model and a Lewis Dot diagram for the element on the board.

1st period Nitrogen



5th period Aluminum